

## Accepted Article

**Title:** Encapsulation and Polymerization of White Phosphorus Inside Single-wall Carbon Nanotubes

**Authors:** Martin Hart, Edward R White, Ji Chen, Catriona M. McGilvery, Chris J. Pickard, Angelos Michaelides, Andrea Sella, Milo Shaffer, and Christoph Günter Salzmann

This manuscript has been accepted after peer review and appears as an Accepted Article online prior to editing, proofing, and formal publication of the final Version of Record (VoR). This work is currently citable by using the Digital Object Identifier (DOI) given below. The VoR will be published online in Early View as soon as possible and may be different to this Accepted Article as a result of editing. Readers should obtain the VoR from the journal website shown below when it is published to ensure accuracy of information. The authors are responsible for the content of this Accepted Article.

**To be cited as:** *Angew. Chem. Int. Ed.* 10.1002/anie.201703585  
*Angew. Chem.* 10.1002/ange.201703585

**Link to VoR:** <http://dx.doi.org/10.1002/anie.201703585>  
<http://dx.doi.org/10.1002/ange.201703585>

# Encapsulation and Polymerization of White Phosphorus Inside Single-wall Carbon Nanotubes

Martin Hart, Edward R. White, Ji Chen, Catriona M. McGilvery, Chris J. Pickard, Angelos Michaelides, Andrea Sella, Milo S. P. Shaffer, and Christoph G. Salzmann\*

**Abstract:** Elemental phosphorus displays an impressive number of allotropes with highly diverse chemical and physical properties. Here, we report that white phosphorus can be filled into single-wall carbon nanotubes (SWCNTs) from the liquid and thereby stabilized against the highly exothermic reaction with atmospheric oxygen. The encapsulated tetraphosphorus molecules were visualized with transmission electron microscopy, but found to convert readily to chain structures inside the SWCNT 'nanoreactors'. The energies of the possible chain structures were determined computationally highlighting a delicate balance between the extent of polymerization and the SWCNT diameter. Experimentally, a single-stranded zig-zag chain of phosphorus atoms was observed which represents the lowest energy structure at small confinement diameters. These one-dimensional chains provide a glimpse into the very first steps of the transformation from white to red phosphorus.

The discovery of phosphorene, a single, two-dimensional sheet of black phosphorus,<sup>[1]</sup> has sparked a renewed interest in the highly diverse chemistry and complex structures of this important main group element. Unlike graphene, phosphorene possesses a direct bandgap that can be tuned by the number of stacked layers, and future applications in electronics and optoelectronics are therefore envisaged.<sup>[1a,1b,2]</sup> But other nanostructured phosphorus allotropes are attracting considerable attention as well. Most recently, blue phosphorus has been grown on gold substrates using molecular beam epitaxial growth<sup>[3]</sup> and a ring-shaped fibrous structure with

alternating P<sub>8</sub> and P<sub>2</sub> structural units has been obtained from a chemical vapor reaction inside multi-wall carbon nanotubes.<sup>[4]</sup> Other allotropes of phosphorus include white phosphorus, which consists of tetrahedral P<sub>4</sub> molecules,<sup>[5]</sup> the amorphous red phosphorus,<sup>[6]</sup> P<sub>4</sub> and violet phosphorus,<sup>[7]</sup> fibrous phosphorus,<sup>[8]</sup> phosphorus nanorods,<sup>[9]</sup> and finally black phosphorus<sup>[6a, 10]</sup> which is the most stable form at ambient pressure.<sup>[11]</sup>

In addition to the current quest for new phosphorus nanomaterials, efforts have been made to stabilize P<sub>4</sub> molecules against the infamous spontaneous combustion with atmospheric oxygen. Recently, P<sub>4</sub> was successfully encapsulated within supramolecular materials<sup>[12]</sup> which may be useful for mitigating dangerous phosphorus spills, or conversely, for controlled release of P<sub>4</sub> into chemical reactions or for heat generation.<sup>[12a]</sup> Single-wall carbon nanotubes (SWCNTs) offer an appealing potential encapsulation environment, with an inert, light-weight, cylindrical pore well-matched to the diameter of P<sub>4</sub> molecules. SWCNTs have been filled with a wide range of materials;<sup>[13]</sup> molecular species can be stabilized sufficiently to allow direct electron microscopy imaging, including of simple chemical reactions.<sup>[14]</sup>

Here we will fill SWCNTs with white phosphorus and explore the stability of the resulting structures. In addition to the experimental approaches, the possible chain structures arising from phase transformations are evaluated using density functional theory (DFT).

Successful filling was achieved by stirring tip-opened CoMoCat SWCNTs,<sup>[15]</sup> with average diameter of ~8.1 Å, in liquid phosphorus at 50°C under inert gas conditions, followed by vacuum sublimation of the excess solid phosphorus at room temperature. The encapsulated P<sub>4</sub> molecules were imaged by high-resolution transmission electron microscopy (HRTEM) (Figure 1a), showing long strings of P<sub>4</sub> beads. At high magnification (Figure 1b), the images are in good agreement with the simulated HRTEM image obtained from the corresponding molecular structure (Figure 1c-d). The presence of phosphorus in the SWCNT samples was also confirmed by X-ray photoelectron spectroscopy (XPS) measurements (Figure 1e); the filled SWCNT sample contains 15 at% phosphorus, 32% of which can be attributed to elemental phosphorus as shown in the inset of Figure 1e. This corresponds to 9.5% elemental phosphorus by weight. Prolonged exposure of the sample to air at room temperature did not change the proportion of elemental phosphorus stabilized within the SWCNTs. The oxidized phosphorus species are therefore assumed to form upon the first exposure of the sample to air; any remaining white phosphorus, either trapped between SWCNT bundles or present at the open tips of the SWCNTs, will oxidize at this point. Additional stirring of the sample in water decreased the content of the phosphorus oxides, but also of elemental phosphorus. Most likely, the oxides passivate the open ends of the SWCNTs; removal simply exposes more elemental phosphorus to the atmosphere.

[\*] M. Hart, Prof. A. Sella, Dr C. G. Salzmann  
Department of Chemistry  
University College London  
20 Gordon Street  
London WC1H 0AJ, United Kingdom  
E-mail: c.salzmann@ucl.ac.uk

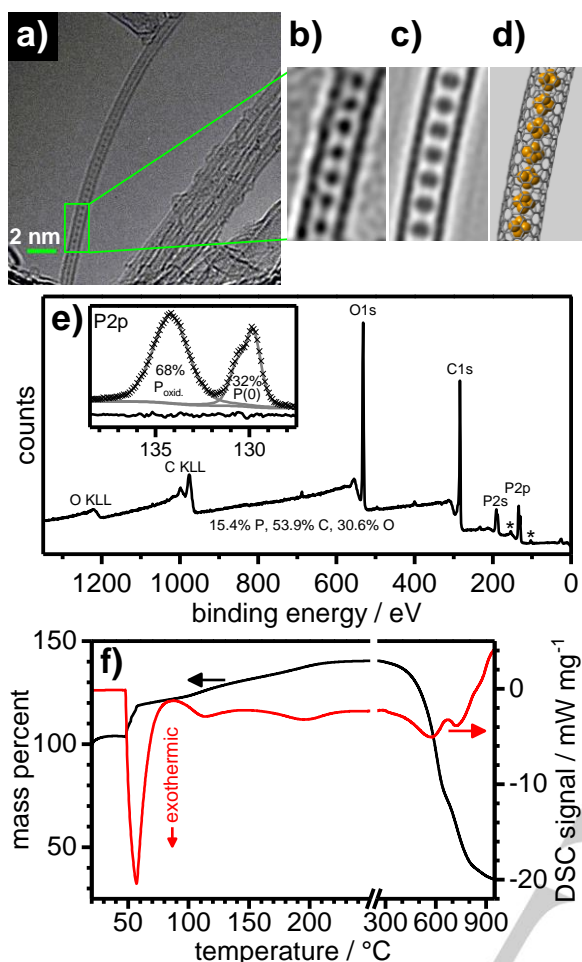
Dr E. R. White, Prof. M. S. P. Shaffer  
Department of Chemistry  
Imperial College London  
Imperial College Road  
London SW7 2AZ, United Kingdom

Dr J. Chen, Prof. A. Michaelides  
Thomas Young Centre, Department of Physics and Astronomy, and  
London Centre for Nanotechnology  
University College London  
Gower Street  
London WC1E 6BT, United Kingdom

Dr C. M. McGilvery,  
Department of Materials  
Imperial College London  
Prince Consort Road  
London SW7 2AZ, United Kingdom

Prof. C. J. Pickard  
Department of Materials Science and Metallurgy  
University of Cambridge  
27 Charles Babbage Road  
Cambridge CB3 0FS, United Kingdom

Supporting information for this article is given via a link at the end of the document.



**Figure 1.** Putting the peas into the pod: SWCNTs filled with  $P_4$  molecules. (a) HRTEM image of a SWCNT filled with a string of  $P_4$  molecules. (b) Noise-filtered HRTEM image of the region highlighted in (a). (c) Simulated HRTEM image and (d) the corresponding atomic structure of  $P_4@SWCNT$ . Panels (b-d) are shown at the same scale. (e) XPS analysis of  $P_4@SWCNT$ . Asterisks indicate a small Si impurity. A high-resolution spectrum of the P2p region is shown in the inset. Spin-orbital splitting into  $2p_{3/2}$  and  $2p_{1/2}$  components can be seen for the XPS feature related to elemental phosphorus. (f) DSC/TGA analysis of  $P_4@SWCNT$ s.

The enthalpy of combustion of  $P_4$  to  $P_4O_{10}$  is remarkably exothermic ( $.2971 \text{ kJ mol}^{-1} P_4$ )<sup>[16]</sup> implying that the new  $P_4@SWCNT$  hybrid may be considered to be a potentially interesting high-energy material. Previous attempts to stabilize  $P_4$  have achieved a maximal loading of 3.5 w%  $P_4$ ,<sup>[12a]</sup> whereas the encapsulation within SWCNTs reported here reached 9.5 w% according to XPS analysis. Assuming a 100% filling yield and complete removal of phosphorus oxide by-products, up to 17 w% could be achieved in principle, as estimated from the available space within the CoMoCat SWCNTs and the van der Waals volume of  $P_4$ .

Differential scanning calorimetry (DSC) data (Figure 1f) shows that the combustion of the encapsulated  $P_4$  can be conducted in a controlled fashion by heating the material in air. The strongest exothermic feature is observed above 50°C which is just above the melting point of white phosphorus at 44.1°C.<sup>[11]</sup>

The two weaker exotherms at 110 and 200°C, respectively, are assigned to the combustion of  $P_4$  that has either undergone conversion towards red phosphorus or been heavily passivated by phosphorus oxides. The exotherms above 300°C are attributed to the burning of the SWCNTs. In line with these assignments, the thermogravimetric analysis (TGA) in Figure 1f shows that the first three exotherms go along with mass increases as the phosphorus converts to phosphorus oxide, whereas the burning of the SWCNTs to give  $CO_2$  gas is accompanied by a mass decrease. Integration over the first three exotherms shows that the combustion of the encapsulated  $P_4$  releases .277 J per milligram of sample. Considering that the combustion enthalpy of pure  $P_4$  is  $23.98 \text{ J mg}^{-1}$ ,<sup>[16]</sup> it follows that the filled SWCNT sample contains 11.6 w% elemental phosphorus which is in good agreement with the value of 9.5 w% obtained from XPS.

Extensive investigation of the new nanomaterial with HRTEM showed that the encapsulated elemental phosphorus is not always present in the form of  $P_4$  molecules. On several occasions, the  $P_4$  molecules had undergone polymerization reactions to form chain structures. We assume that this is either aided by the exposure to the electron beam, the result of previous thermal treatments, or promoted by the presence of the SWCNTs. To our surprise, it has not been possible to observe the highly Raman-active breathing mode<sup>[17]</sup> of the encapsulated  $P_4$  molecules (Supporting Information, Figure S1). This failure may lend support to the idea that both the electron beam or the incident photons of the Raman laser may drive the polymerization of the encapsulated  $P_4$  molecules. It is well known that the conversion of bulk white to red phosphorus can be induced by exposure to visible light.<sup>[11]</sup>

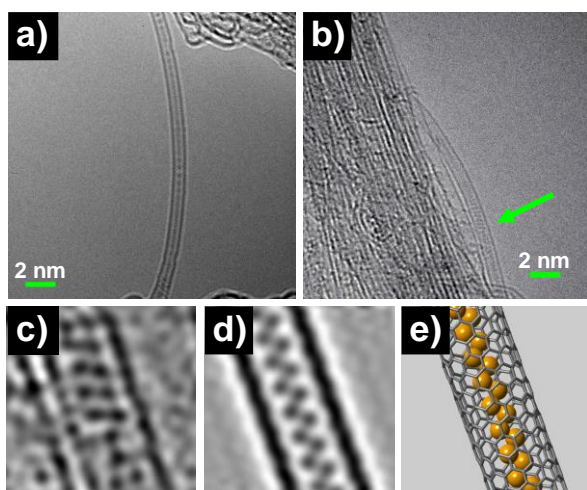
The HRTEM image in Figure 2a shows the coexistence of  $P_4$  molecules and linear chain structures in the fashion of a } } [ • & # 247; A & # 247; [ ! • ^ A & # 247; a ^ q In general, the atomic structure of the chains was difficult to resolve with HRTEM. Considering that the high-temperature U phase of white phosphorus is a plastic crystal with dynamic reorientations of the  $P_4$  molecules,<sup>[5b]</sup> it is likely that both the encapsulated  $P_4$  molecules, as well as the associated chain structures, also display pronounced dynamic behavior which makes it difficult to achieve atomic resolution with HRTEM. However, analysis of a successful image (Figure 2b) with the aid of a simulated HRTEM image is consistent with a single-stranded zig-zag chain of phosphorus atoms (Figure 2c. e). Furthermore, the line features shown in Figure 2a are consistent with the simulated HRTEM image of a rotationally-averaged single-stranded zig-zag chain (Supporting Information, Figure S2). A tendency for the formation of chain structures inside SWCNTs has also been observed for other elements including iodine,<sup>[13a]</sup> sulphur,<sup>[13c]</sup> selenium,<sup>[13d]</sup> and even carbon itself in the form of encapsulated carbyne<sup>[13e]</sup> or chains of nanodiamonds<sup>[13f]</sup>.

Insights into charge transfer processes between phosphorus and carbon nanomaterials can be obtained from Raman spectroscopy.<sup>[4,18]</sup> As demonstrated in ref. [4], the presence of phosphorus oxides leads to hole-doping and consequently to a blue-shift of the G band of multi-wall carbon nanotubes, whereas the electron-doping from the lone pairs of elemental phosphorus results in small red shifts. Consistent with this, a SWCNT sample containing predominately oxidized



## COMMUNICATION

phosphorus species according to XPS analysis showed a pronounced blue shift of the G band of  $11\text{ cm}^{-1}$ .<sup>[18]</sup> Depending on the excitation wavelength, the G bands of our  $\text{P}_4$ -filled SWCNTs were either slightly red-shifted ( $\sim 1.5\text{ cm}^{-1}$ ) or unchanged within the error of the measurement compared to the empty SWCNTs (Supporting Information, Figure S3). This implies that the phosphorus oxides present in our samples do not significantly influence the electronic properties of the SWCNTs.

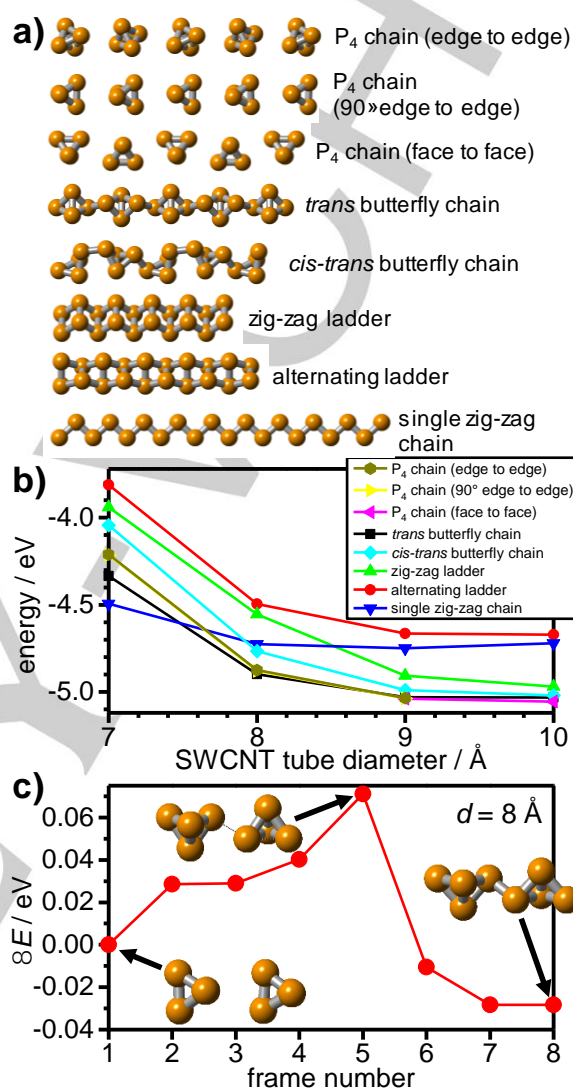


**Figure 2.** Polymerization of  $\text{P}_4$  molecules inside SWCNTs. (a+b) HRTEM images of SWCNTs containing polymeric phosphorus structures. (c) Noise-filtered HRTEM image of the region highlighted in (b). (d) Simulated HRTEM image and (e) the corresponding atomic structure of a single zig-zag chain@SWCNT. Panels (c-e) are shown at the same scale.

In addition to the experiments, ab initio random structure searches (AIRSS)<sup>[19]</sup> were used to identify the possible chain structures and the most relevant structures are shown in Figure 3a. Starting from different packings of  $\text{P}_4$  molecules, a single bond formation between neighboring  $\text{P}_4$  molecules gives chains [  $\text{P}_4$  chain (edge to edge),  $\text{P}_4$  chain (90° edge to edge),  $\text{P}_4$  chain (face to face) ] -atom bridges can be either *trans* or *cis-trans* with respect to the newly formed P-P bonds. Pauling proposed that these structures are the first step in the phase transition from white to red phosphorus.<sup>[20]</sup> Further rearrangements of the butterfly units then lead to ladder structures. The zig-zag ladder is interesting as this structural motive is found in the fibrous and Hittorf polymorphs<sup>[7a,8]</sup> as well as in the recently discovered ring-shaped phosphorus.<sup>[4]</sup> Yet, in those cases, the zig-zag ladder motifs are stabilized with  $\text{P}_4$  molecules which cannot be accommodated within the confinement of SWCNTs. Finally, the single-stranded zig-zag chain can be considered to be the final state of a polymerization reaction achieved by successive deconstruction of the initial  $\text{P}_4$  tetrahedra and it is the only structure in Figure 3a with free valencies.

Figure 3b shows the DFT energies of the various chain structures including contributions from binding and confinement as a function of the SWCNT diameter. Generally speaking, the energy differences between the various chain structures are quite small for a given SWCNT diameter. At larger diameters, the free  $\text{P}_4$  molecules are most stable. Yet, the energy difference

between the free molecules and in particular the *trans* butterfly chain is very small. Interestingly, towards smaller SWCNT diameters, the single zig-zag chain becomes most stable.



**Figure 3.** DFT analysis of filling and polymerizing  $\text{P}_4$  molecules inside SWCNTs. (a) Geometry-optimized one-dimensional phosphorus structures. As drawn, all chains contain 20 P atoms. (b) Energies of the chain structures shown in (a) as a function of the SWCNT diameter. (c) Energy differences relative to the starting structure along a computationally optimized pathway from edge-to-edge packed  $\text{P}_4$  molecules to the *trans* butterfly chain determined using the climbing-image nudged-elastic-band (cNEB) method and an 8 Å diameter SWCNT. The energies in (b) and (c) are given per P atom.

Using nudged-elastic-band DFT calculations<sup>[21]</sup> the activation energy for the reaction from  $\text{P}_4$  molecules to the *trans* butterfly chain was determined as  $0.07\text{ eV mol}^{-1}$  (Figure 3c). Considering that  $kT$  has a value of  $0.0257\text{ eV}$  at room temperature, it is now understandable that confined  $\text{P}_4$  molecules are likely to undergo transformation readily. Furthermore, considering the small energy differences between the various chain structures, it seems reasonable to assume that the polymerization reactions are, at least to some extent,

## COMMUNICATION

reversible within the confinement of SWCNTs. Due to the different space requirements of the various chain structures (Figure 3a), it is even conceivable that different extents of lateral pressures could affect the local equilibria. Considering that complicated chemical equilibria may exist between the various one-dimensional structures, it is important to stress that our TEM observations may not necessarily represent the lowest energy structures for a given SWCNT diameter.

In summary, SWCNTs have been filled with P<sub>4</sub> molecules at filling yields of up to ~10 w%. The encapsulated white phosphorus is stable at room temperature in air. Yet, upon heating above 50°C, the highly exothermic reaction with oxygen takes place. This small temperature window makes our new composite material well-suited for applications where controlled releases of heat are required. A tendency for the confined P<sub>4</sub> molecules to polymerize has been observed. According to DFT calculations, a range of possible chain structures is expected to exist including a single zig-zag chain of phosphorus atoms which has been observed in HRTEM. These polymerizations are thought to represent the very first steps in the transformation from white to red phosphorus. Future explorations of their electronic properties and chemical interactions will be of great interest. Their encapsulation provides an exciting opportunity to isolate and study highly reactive intermediates in the phase behavior of phosphorus.

## Acknowledgements

We thank the Advanced Characterization of Materials CDT for a studentship (MH), the Royal Society for a University Research Fellowship (CGS) and Wolfson Research Merit Awards (CJP & AM), the European Research Council under the EU FP7 Framework Programme (FP/2007-2013) / ERC Grant Agreement number 616121 (JC & AM, Heterolce project). Computational support from the UK national high-performance computing service, ARCHER, was obtained via the UKCP consortium (EPSRC grants EP/P022596/1 and EP/P022103/1).

**Keywords:** Phosphorus Allotropy Nanotubes Polymerization Density functional calculations

- [1] a) H. Liu, A. T. Neal, Z. Zhu, Z. Luo, X. Xu, D. Tománek, P. D. Ye, *ACS Nano* **2014**, *8*, 4033-4041; b) L. Li, Y. Yu, G. J. Ye, Q. Ge, X. Ou, H. Wu, D. Feng, X. H. Chen, Y. Zhang, *Nat. Nano.* **2014**, *9*, 372-377; c) S. P. Koenig, R. A. Doganov, H. Schmidt, A. H. C. Neto, B. Özyilmaz, *Appl. Phys. Lett.* **2014**, *104*, 103106; d) C.-G. Andres, V. Leonardo, P. Elsa, O. I. Joshua, K. L. Narasimha-Acharya, I. B. Sofya, J. G. Dirk, B. Michele, A. S. Gary, J. V. Alvarez, W. Z. Henny, J. J. Palacios, S. J. v. d. Z. Herre, *2D Materials* **2014**, *1*, 025001.
- [2] a) F. Xia, H. Wang, Y. Jia, *Nat. Comm.* **2014**, *5*, 4458; b) Z. Guo, H. Zhang, S. Lu, Z. Wang, S. Tang, J. Shao, Z. Sun, H. Xie, H. Wang, X.-F. Yu, P. K. Chu, *Adv. Funct. Mater.* **2015**, *25*, 6996-7002.
- [3] J. L. Zhang, S. Zhao, C. Han, Z. Wang, S. Zhong, S. Sun, R. Guo, X. Zhou, C. D. Gu, K. D. Yuan, Z. Li, W. Chen, *Nano Lett.* **2016**, *16*, 4903-4908.
- [4] J. Zhang, D. Zhao, D. Xiao, C. Ma, H. Du, X. Li, L. Zhang, J. Huang, H. Huang, C.-L. Jia, D. Tománek, C. Niu, *Angew. Chem. Int. Ed.* **2017**, *56*, 1850-1854; *Angew. Chem.* **2017**, *129*, 1876-1880.
- [5] a) D. E. C. Corbridge, E. J. Lowe, *Nature* **1952**, *170*, 629-629; b) A. Simon, H. Borrmann, J. Horak, *Chem. Ber.* **1997**, *130*, 1235-1240.
- [6] a) R. Hultgren, N. S. Gingrich, B. E. Warren, *J. Chem. Phys.* **1935**, *3*, 351-355; b) J. M. Zaugg, A. K. Soper, S. M. Clark, *Nat. Mater.* **2008**, *7*, 890-899.
- [7] a) W. Hittorf, *Ann. Phys. (Berlin)* **1865**, *202*, 193-228; b) G. Schusteritsch, M. Uhrin, C. J. Pickard, *Nano Lett.* **2016**, *16*, 2975-2980.
- [8] a) M. Ruck, D. Hoppe, B. Wahl, P. Simon, Y. Wang, G. Seifert, *Angew. Chem. Int. Ed.* **2005**, *44*, 7616-7619; *Angew. Chem.* **2005**, *117*, 7788-7792; b) A. Pfizner, M. F. Bräu, J. Zweck, G. Bruncklaus, H. Eckert, *Angew. Chem. Int. Ed.* **2004**, *116*, 4324-4327; *Angew. Chem.* **2004**, *116*, 4324-4327.
- [9] R. A. L. Winchester, M. Whitby, M. S. P. Shaffer, *Angew. Chem. Int. Ed.* **2009**, *48*, 3616-3621; *Angew. Chem.* **2009**, *121*, 3670-3675.
- [10] A. H. Woomey, T. W. Farnsworth, J. Hu, R. A. Wells, C. L. Donley, S. C. Warren, *ACS Nano* **2015**, *9*, 8869-8884.
- [11] M. E. Schlesinger, *Chem. Rev.* **2002**, *102*, 4267-4302.
- [12] a) P. Mal, B. Breiner, K. Rissanen, J. R. Nitschke, *Science* **2009**, *324*, 1697-1699; b) C. Schwarzmaier, A. Schindler, C. Heindl, S. Scheuermayer, E. V. Peresyphkina, A. V. Virovets, M. Neumeier, R. Gschwind, M. Scheer, *Angew. Chem. Int. Ed.* **2013**, *52*, 10896-10899; *Angew. Chem.* **2013**, *125*, 11097-11100; c) D. Yang, J. Zhao, L. Yu, X. Lin, W. Zhang, H. Ma, A. Gogoll, Z. Zhang, Y. Wang, X.-J. Yang, B. Wu *J. Am. Chem. Soc.* **2017**, *139*, 5946-5951.
- [13] a) X. Fan, E. C. Dickey, P. C. Eklund, K. A. Williams, L. Grigorian, R. Buczko, S. T. Pantelides, S. J. Pennycook, *Phys. Rev. Lett.* **2000**, *84*, 4621-4624; b) R. R. Meyer, J. Sloan, R. E. Dunin-Borkowski, A. I. Kirkland, M. C. Novotny, S. R. Bailey, J. L. Hutchison, M. L. H. Green, *Science* **2000**, *289*, 1324-1326; c) T. Fujimori, A. Morelos-Gómez, Z. Zhu, H. Muramatsu, R. Futamura, K. Urita, M. Terrones, T. Hayashi, M. Endo, S. Young Hong, Y. Chul Choi, D. Tománek, K. Kaneko, *Nat. Comm.* **2013**, *4*, 2162; d) T. Fujimori, R. B. dos Santos, T. Hayashi, M. Endo, K. Kaneko, D. Tománek, *ACS Nano* **2013**, *7*, 5607-5613; e) L. Shi, P. Rohringer, K. Suenaga, Y. Niimi, J. Kotakoski, J. C. Meyer, H. Peterlik, M. Wanko, S. Cahangirov, A. Rubio, Z. J. Lapin, L. Novotny, P. Ayala, T. Pichler, *Nat. Mater.* **2016**, *15*, 634-639; f) Y. Nakanishi, H. Omachi, N. A. Fokina, P. R. Schreiner, R. Kitaura, J. E. P. Dahl, R. M. K. Carlson, H. Shinohara, *Angew. Chem. Int. Ed.* **2015**, *54*, 10802-10806; *Angew. Chem.* **2015**, *127*, 10952-10956.
- [14] A. Botos, J. Biskupek, T. W. Chamberlain, G. A. Rance, C. T. Stoppiello, J. Sloan, Z. Liu, K. Suenaga, U. Kaiser, A. N. Khlobystov, *J. Am. Chem. Soc.* **2016**, *138*, 8175-8183.
- [15] S. M. Bachilo, L. Balzano, J. E. Herrera, F. Pompeo, D. E. Resasco, R. B. Weisman, *J. Am. Chem. Soc.* **2003**, *125*, 11186-11187.
- [16] W. S. Holmes, *Trans. Faraday Soc.* **1962**, *58*, 1916-1925.
- [17] I. R. Beattie, G. A. Ozin, R. O. Perry, *J. Chem. Soc. A* **1970**, 2071-2074.
- [18] a) J. Xu, L. Guan, *J. Phys. Chem. C* **2009**, *113*, 15099-15101; b) J. Xu, L. Guan, *RSC Adv.* **2013**, *3*, 5577-5582.
- [19] C. J. Pickard, R. J. Needs, *J. Phys.: Condens. Matter* **2011**, *23*, 053201.
- [20] L. Pauling, M. Simonetta, *J. Chem. Phys.* **1952**, *20*, 29-34.
- [21] G. Henkelman, B. P. Uberuaga, H. Jónsson, *J. Chem. Phys.* **2000**, *113*, 9901-9904.

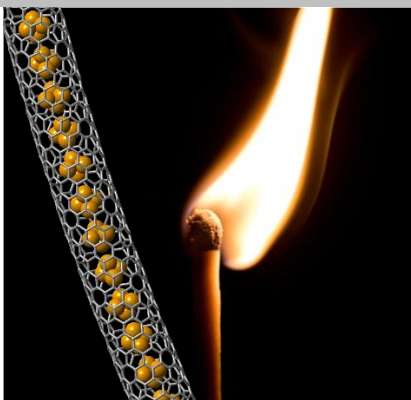
## COMMUNICATION

## Entry for the Table of Contents

Layout 1:

## COMMUNICATION

Tetrahedral  $P_4$  molecules can be assembled like peas in a pod inside single wall carbon nanotubes to give air-stable constructs. The encapsulated molecules show a tendency to polymerise giving rise to range of possible chain structures including a zig-zag chain of individual atoms. These one-dimensional allotropes provide the first direct glimpse into the transformation from white to red phosphorus.



*M. Hart, E. R. White, J. Chen, C. M. McGilvery, C. J. Pickard, A. Michaelides, A. Sella, M. S. P. Shaffer and C. G. Salzmann\**

**Page No. – Page No.**

**Encapsulation and polymerization of white phosphorus inside single-wall carbon nanotubes**