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July 2012, page 10
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The article by Thorne Lay and Hiroo Kanamori (10.1063/PT.3.1619) is an excellent review of the 1994 Chilean earthquake. The authors estimate that the earthquake released approximately five times as much energy as the 1945 Nagasaki atomic bombing. This is a significant finding, especially since the 1945 atomic bombing is often cited as the most powerful man-made explosion in history. The authors' calculations are based on the seismic energy released by the earthquake, which is a much more direct measure of the energy released than the energy released by the atomic bombing. The authors also note that the seismic energy released by the earthquake is comparable to the energy released by the 1945 atomic bombing. This is a significant finding, especially since the 1945 atomic bombing is often cited as the most powerful man-made explosion in history. The authors' calculations are based on the seismic energy released by the earthquake, which is a much more direct measure of the energy released than the energy released by the atomic bombing. The authors also note that the seismic energy released by the earthquake is comparable to the energy released by the 1945 atomic bombing.

By the act of hitting a ball with a bat, one calculates the force energy to deliver the ball to its new location, but one must also take into account that the ball extended its energy release to that which became struck by the ball as its momentum ceased and passed energy to the struck team. Therefore the parameters of the damage extend into the future when the received energy to that pushed upon, later becomes released in a new event. Perhaps calculations of one added that in while another's calculations did not. E.M.C.
Written by Edgar Mocarvill, 14 July 2012 19:59

Temperature dependent partition functions and equilibrium constant for HCN and HNC

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Ab initio vibration-rotation energy levels are summed to estimate a partition function for the total HCN system. By assigning individual levels to HCN and HNC, separate partition functions are obtained for the isomers. These are used to give a temperature dependent equilibrium constant which suggests that at temperatures typical of cool carbon stars, about 20% of the HCN system is actually HNC. Errors in the partition functions and equilibrium constant are estimated. © 2002 American Institute of Physics. [DOI: 10.1063/1.1521131]

I. INTRODUCTION

The HCN and HNC linear molecules provide an important and well-studied prototype of an isomerizing system. In particular these molecules are probably the only chemically bound isomers amenable to full, accurate *ab initio* study not only of their electronic potential energy surfaces, see, for example, Refs. 1–4, but also their nuclear motion states up to and including those “delocalized” states which link the two linear minima.^{5–10}

The HCN system is important in its own right. It is a trace constituent of our own atmosphere and is widely seen in space. In contrast, the high abundance of the metastable HNC isomer which is frequently observed in cold, dense molecular clouds^{11,12} remains something of a mystery. Such non-LTE (local thermodynamic equilibrium) behavior is outside the scope of the present study. Of more relevance is the presence of HCN in the atmospheres of cool carbon stars where it provides a major source of opacity.^{13,14} The atmospheres of cool stars are LTE environments and HCN is widely observed at temperatures of 2000–3000 K. As we show below, at such temperatures there should also be a significant amount of HNC present in the atmospheres.

There are various methods of computing molecular partition functions. One that we have exploited successfully for water, up to temperatures of over 5000 K,^{15,16} is the explicit summation of energy levels obtained from variational nuclear motion calculations. No calculation has yet obtained *all* the bound rotation-vibration energy levels for a chemically bound triatomic. This means that in practice, in order to obtain convergence at higher temperatures, explicit summations over calculated energy levels must be augmented with estimates for higher-lying levels not otherwise included. This is the strategy that we employ here.

In this work we compute three separate temperature-dependent partition functions: one for the whole HCN system, one for HCN and one for HNC. Clearly, to obtain the last two partition functions it is necessary to identify individual energy levels as belonging to either HCN or HNC. We then use these two partition functions to estimate the HNC \leftrightarrow HCN equilibrium constant as a function of temperature. Thermodynamic data for HCN, computed from its partition

function, is available in JANAF¹⁷ and fits to the JANAF partition function itself have been presented by Irwin^{18,19} over an extended temperature range. However the spectroscopic data used by JANAF and hence Irwin to generate the partition function is both old and very limited. Whereas there are some data on the HCN partition function, there appears to be no data available for HNC. There is only one previous estimate for the equilibrium constant, an unreliable empirically-determined one by Maki and Sams²⁰ which gave an energy separation between the isomers significantly lower than all previous and subsequent determinations.

As our calculations use energy levels from first principles, variational nuclear motion calculations^{21,22} based on the use of an *ab initio* potential energy surface,⁴ they themselves can be considered to be *ab initio*.

II. METHOD

The present calculations are based on the energy levels calculated by Harris *et al.*²² who used them and the associated wave functions to give a comprehensive line-list of transitions for the HCN/HNC system. These energy levels were found by solving the bound state nuclear motion for the VQZANO+ potential of Van Mourik *et al.*⁴ This surface was obtained by combining *ab initio* electronic structure calculation for the isomerizing HCN/HNC system at several levels of theory including explicit allowance for electronic relativistic effects and the adiabatic correction to the Born Oppenheimer approximation. Although not as accurate for the known energy levels as surfaces that have been empirically adjusted to spectroscopic data,²³ the VQZANO+ is the best available *ab initio* potential for the system. With this surface the HNC metastable isomer lies at $\Delta E_{\text{HNC}} = 5185.6 \text{ cm}^{-1}$ above the HCN isomer, once allowance is made for zero point energy corrections. This prediction is in agreement with the experimental results of Pau and Hehre,²⁴ who obtained $5180 \pm 700 \text{ cm}^{-1}$, rather than $3600 \pm 400 \text{ cm}^{-1}$ given by Maki and Sams.²⁰ Pau and Hehre's result is supported by a recent reanalysis by Wenthold.²⁵

The VQZANO+ potential gives a barrier to isomerization which is $16\,798 \text{ cm}^{-1}$ above the HCN minimum.⁴ Delocalized states, which cross the barrier and therefore cannot

be uniquely associated with a particular isomer, occur at energies above $15\,000\text{ cm}^{-1}$ relative to the HCN ground state, see Bowman *et al.*¹⁰ for a detailed discussion of this.

Harris *et al.* used the exact kinetic energy operator nuclear motion program suite DVR3D²⁶ to obtain rotation-vibration energy levels and corresponding wave functions. They considered all levels of the system with rotational angular momentum, J , up to 60 which lay up to $18\,000\text{ cm}^{-1}$ above the ground state of HCN, which is taken below as our zero of energy. Details of the calculation procedure used are given in Ref. 22 and the levels can be obtained by either anonymous ftp from ftp.tampa.phys.ucl.ac.uk and changing to the directory /pub/astrodata/HCN, or via the web page <http://www.tampa.phys.ucl.ac.uk> and following the links for "Astrodata and Molecular data" then "Astrodata" then "HCN."

Comparisons with experiment^{4,21} suggest that the vibrational spacings agree with experiment to within a few cm^{-1} , or 0.15%, in most cases with the worst case being the HNC v_1 fundamental mode where the error rises to 0.4%. The calculations are particularly good at reproducing the observed rotational structure of the system;²¹ the errors for the rotational levels are therefore less than the vibrational ones. These systematic errors dominate the errors in our partition functions at low temperature.

Altogether, Harris *et al.*²² computed over 168 000 energy levels for the HCN/HNC system. At temperatures where this number of levels is sufficient to converge the summation, computation of the partition function for the total system can be achieved by a straightforward summation over these levels. However, to obtain separate partition functions for HCN and HNC is more difficult. This is because the energy levels obtained by Harris *et al.* are not labeled as to which isomer they belong to. It is therefore necessary, at least in principle, to assign each energy level to an isomer. For 168 000 levels this is a formidable task and we have therefore made some simplifying assumptions discussed below which mean that in practice we have explicitly assigned only approximately 10 000 levels to an isomer. The validity of several of these assumptions can be checked by showing that they result in a negligible error in the equivalent partition function for the total system.

Both HCN and HNC are linear molecules. This means that for states with $J>0$ and $l>0$ there are two, nearly degenerate states, conventionally designated as e and f . For $J=0$ only the e state exists. Since the splitting between e and f states is very small we have only analyzed the e states and doubly weighted those levels with $l>0$.

To analyze the energy level data, all data for e states with $J\leq 38$ was read into a spreadsheet. These data were then aligned so that states with the same J were all in the same column and states with a common set of vibrational quantum numbers and belonging to same isomer were in the same row. As there were many state crossings in the original, purely energy-ordered, data many realignments were necessary. States were aligned by computing the rigid rotor quantum number, B_v , using a formula based on three neighboring states. In the course of sorting the data it was also possible to assign the vibrational angular momentum quantum number,

ℓ , for each state since energy levels only exist for which $\ell\leq J$. Thus for $J=0$ there are only states with $\ell=0$, new vibrational states appearing for $J=1$ have $\ell=1$ and so forth. For $J=14$ to 38, 302 vibrational states were aligned, spanning energies up to $12\,500\text{ cm}^{-1}$. Because vibrational states with ℓ up to 12 were analyzed, there are somewhat fewer states for lower J values.

Having aligned the data, values for the rotational constants B_v and D_v were obtained for each vibrational series using the fourth-order rotational term value formula

$$E_r = B_v[J(J+1) - \ell^2] - D_v[J(J+1) - \ell^2]^2. \quad (1)$$

HCN levels could readily be associated with those levels for which $B_v\leq 1.51\text{ cm}^{-1}$ and HNC with B_v values above this value. However some HCN states with relatively high values of D_v were found to have B_v values somewhat above 1.51 cm^{-1} . By analyzing the constants for each case, isomer assignments were made to each of the vibrational states. These assignments, with only one exception, were found to agree with the previous analysis of Harris *et al.*²¹ who made incomplete but more extensive isomer assignments on the basis of dipole transition intensities. A third method of making possible isomer assignments, involving the sign of the permanent dipole moment,¹⁰ was not tested here.

At laboratory temperatures the energy levels considered above are sufficient to converge the partition function sums, but for higher temperatures it is necessary to consider the contribution from both higher rotational and higher vibrational states. In practice, because of the $(2J+1)$ weighting in the partition function sum, the inclusion of higher rotational states is the more important at temperatures up to 4000 K (at higher temperatures, the omission of higher vibrational states becomes increasingly significant). States with $J>38$ were included by extrapolating using Eq. (1) and the B_v and D_v constants whose determination is discussed above, with the exception that in the few cases where a negative value of D_v was obtained $D_v=0$ was used instead. States up to $J=91$ were explicitly included in our final summation and the contribution from higher J 's (up to 150) was approximated by assuming the contribution to $Q(T)$ from successive J 's followed a geometric progression with a temperature dependent common multiple. The geometric factor, which includes the degeneracy factor, rose from 0.691 at $T=1000\text{ K}$ to 0.912 at $T=4000\text{ K}$, numbers that are consistent with the pattern for lower J 's. At the temperatures considered here, the contribution to Q from $J>150$ was found to be negligible.

Tests using an extra term, H_v , obtained using a sixth-order fit gave very similar results. The lower-order fit was preferred for reasons of numerical stability. In particular B_v was found to always be at least 10^5 greater than D_v which is significantly bigger than the values of J^2 considered.

Below 1000 K the 302 vibrational states explicitly analyzed above are sufficient to give 99.99% of the partition sums. However at 3000 K only about 87% of the partition function arises from these lower vibrational states. Harris *et al.*²¹ assigned approximately 70% of the energy levels between the 302 vibrational levels assigned here and the $18\,000\text{ cm}^{-1}$ cutoff to either HCN, HNC or delocalized

states on the basis of transition intensities. At higher energies, these assignments were subject to increasing uncertainty, nevertheless, Harris *et al.*'s assignments were used above the 302 energy level and the contribution to the partition function sums from the unassigned higher vibrational states was allocated to each isomer by assuming that the gaps in Harris's assignments had no bias to one or the other isotope. Analysis showed that approximately 40% of the higher levels are HCN and 60% HNC. The higher proportion of HNC levels at high energies is due to the much broader potential well about the HNC minimum which leads to a higher density of states. Tests suggest that this assumption leads a possible 5% error in the equilibrium constant K at the highest temperatures considered.

III. RESULTS

Table I presents our calculated partition function for the whole HCN system, $Q(\text{Total})$, the partition function for HCN, $Q(\text{HCN})$, that for HNC, $\tilde{Q}(\text{HNC})$, and the partition function for the delocalized states i.e., those that cannot be assigned to only one isomer, $\tilde{Q}(\text{Deloc})$. The error estimates in Table I are derived from three possible sources of error (a) the underlying error in the *ab initio* data, (b) errors due to the approximate treatment of levels with $J > 38$ and (c) errors due to possible misidentification of isomer levels. At low temperature, (a) is the dominant source of error. For $T \geq 2500$ K there is a further source of error due to the omission of high-lying vibrational states from our model. This will be discussed below.

All partition functions in Table I are expressed using the $J=0$ state of HCN as the energy zero. Of course this is not the standard form for the HNC partition function, $Q(\text{HNC})$, which should be expressed relative to the lowest level of HNC. The two partition functions are related by the expression

$$Q(\text{HNC}) = \exp\left(\frac{\Delta E_{\text{HNC}}}{kT}\right) \tilde{Q}(\text{HNC}), \quad (2)$$

where in this work ΔE_{HNC} is taken to be 5185.6 cm^{-1} . Figure 1 plots the partition functions as a function of temperature showing that $\tilde{Q}(\text{HNC})$ forms a larger proportion of $Q(\text{Total})$ as the temperature increases. Of course this is to be expected, but the effect is increased by the fact that the HNC side of the potential is much broader than the HCN side. This means that HNC supports a higher density of states meaning that actually $Q(\text{HNC})$ is consistently bigger than $Q(\text{HCN})$.

Figure 1 compares our partition functions with the most recent fit by Irwin¹⁹ to the data in JANAF.¹⁷ JANAF is entirely based on HCN data and so, presumably, Irwin's fit should correspond to $Q(\text{HCN})$; in practice it lies intermediate between $Q(\text{HCN})$ and $Q(\text{Total})$.

Maki *et al.*²⁷ give room temperature values for the $Q(\text{HCN})$ as 148.72 at 296 K and 149.94 at 298 K. These values are consistent with our calculations for $Q(\text{HCN})$ which are 148.53 ± 0.54 and 149.75 ± 0.54 at the same temperatures. The corresponding values for $Q(\text{HNC})$ are 169.48 and 171.22, respectively, also with an error of ± 0.54 .

TABLE I. Partition functions of the total HCN system and its component molecules as a function of temperature. The % error applies to both $Q(\text{Total})$ and $Q(\text{HCN})$.

T/K	$Q(\text{Total})$	$Q(\text{HCN})$	$\tilde{Q}(\text{HNC})$	$\tilde{Q}(\text{Deloc})$	% Error
150	71.03	71.03	0	0	0.4
200	95.49	95.49	0	0	0.4
250	121.81	121.81	0	0	0.4
300	150.99	150.99	0	0	0.4
350	183.79	183.79	0	0	0.4
400	220.87	220.87	0	0	0.3
450	262.79	262.79	0	0	0.3
500	310.07	310.07	0	0	0.3
550	363.3	363.3	0	0	0.3
600	422.9	422.9	0	0	0.3
650	489.6	489.6	0	0	0.3
700	564.0	564.0	0	0	0.3
750	646.8	646.7	0	0	0.2
800	738.6	738.5	0.1	0	0.2
850	840.3	840.1	0.2	0	0.2
900	952.6	952.2	0.4	0	0.2
950	1076.5	1075.8	0.7	0	0.2
1000	1212.8	1211.7	1.2	0	0.2
1100	1527.2	1524.2	2.9	0	0.2
1200	1904.5	1897.9	6.6	0	0.2
1300	2355.1	2341.6	13.5	0	0.2
1400	2890.6	2865.0	25.5	0.1	0.2
1500	3523.9	3478.6	45.1	0.2	0.2
1600	4270	4193	76	0.5	0.2
1700	5145	5021	122	1	0.3
1800	6166	5975	189	2	0.3
1900	7355	7067	284	4	0.3
2000	8731	8311	414	7	0.3
2100	10 319	9720	588	12	0.3
2200	12 143	11 308	815	20	0.4
2300	14 228	13 089	1108	31	0.4
2400	16 600	15 077	1477	47	0.5
2500	19 288	17 284	1935	68	0.5
2600	22 317	19 725	2495	97	0.6
2700	25 717	22 410	3172	135	0.6
2800	29 514	25 353	3977	184	0.7
2900	33 734	28 564	4926	244	0.8
3000	38 405	32 054	6031	319	0.8
3100	43 550	35 833	7307	410	0.9
3200	49 194	39 909	8766	519	1.0
3300	55 358	44 291	10 420	647	1.0
3400	62 063	48 986	12 281	796	1.1
3500	69 329	53 999	14 361	969	1.2
3600	77 172	59 337	16 668	1167	1.3
3700	85 608	65 004	19 213	1391	1.3
3800	94 650	71 003	22 004	1643	1.4
3900	104 311	77 338	25 048	1925	1.5
4000	114 599	84 009	28 352	2238	1.6

Since partition functions are usually required at a range of rather precise temperatures we have fitted our data to a standard polynomial form:

$$\log_{10} Q(T) = \sum_{i=0}^4 a_i (\log_{10} T)^i. \quad (3)$$

Constants for $Q(\text{Total})$, $Q(\text{HCN})$ and $Q(\text{HNC})$ are given in Table II. These constants reproduce the partition functions up to 4000 K to within 0.2% or better.

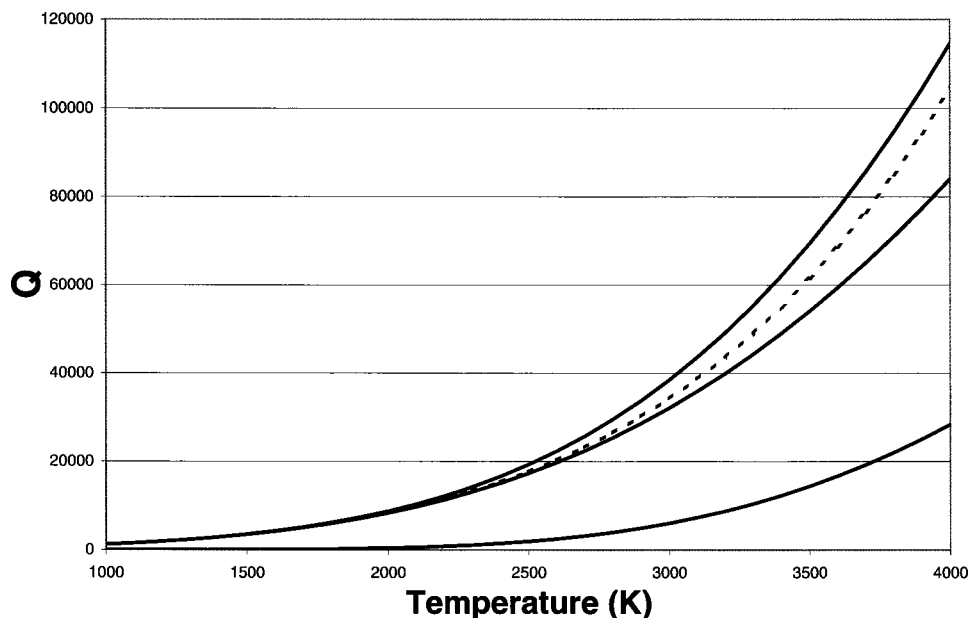


FIG. 1. Partition functions as a function of temperature. Solid curves: this work $Q(\text{Total})$ highest, $Q(\text{HCN})$ middle curve and $Q(\text{HNC})$ lowest. The dashed curve is the HCN partition function given by Irwin (Ref. 19).

The partition functions for HCN and HNC can be used to obtain a temperature-dependent estimate of the equilibrium constant, K , for the simple isomerization reaction



using the formula

$$K = \frac{\tilde{Q}(\text{HNC})}{Q(\text{HCN})} = \frac{Q(\text{HNC})}{Q(\text{HCN})} \times \exp\left(-\frac{\Delta E_{\text{HNC}}}{kT}\right). \quad (5)$$

Calculated values for K are given in Table III which suggests that at temperatures above 1000 K there should be appreciable quantities of HNC present in any thermalized sample. Indeed at the temperature of the atmosphere of cool carbon stars, up to 20% of the HCN system should be in the form of HNC. As the absorption bands HNC are generally stronger than those of HCN,^{21,22} HNC should therefore provide an important source of absorption at these temperatures. Furthermore the ratio of HNC/HCN has the potential to provide a useful temperature diagnostic.

In considering the results presented here it is important to try and assess the errors associated with what is, after all, a completely *ab initio* theoretical treatment of the problem. In such treatments the errors are essentially all systematic which makes the giving of meaningful error bars difficult. These systematic errors essentially arise from two sources: errors in the energy levels used and errors introduced because the data set of levels used is incomplete.

The use of incomplete data necessarily leads to an underestimate of the associated partition functions and the associated error increases with T . Given that we allow, albeit

TABLE III. Equilibrium constant, K , for $\text{HCN} \leftrightarrow \text{HNC}$ as a function of temperature. The percentage of HNC for a thermalized sample is also given. (Numbers in parenthesis denote powers of ten.)

T/K	$K(T)$	% Error	% HNC
300	1(-11)	0.8	2(-9)
400	1(-8)	0.6	1(-6)
500	4(-7)	0.6	4(-5)
600	6(-6)	0.6	6(-4)
700	4(-5)	0.6	4(-3)
800	0.0001	0.4	0.014
900	0.0004	0.4	0.040
1000	0.0010	0.4	0.095
1100	0.0019	0.4	0.193
1200	0.0035	0.4	0.347
1300	0.0058	0.4	0.573
1400	0.0089	0.4	0.881
1500	0.0130	0.4	1.28
1600	0.0181	0.4	1.78
1700	0.0243	0.6	2.38
1800	0.0317	0.6	3.07
1900	0.0402	0.6	3.86
2000	0.0498	0.6	4.74
2100	0.0605	0.6	5.69
2200	0.0721	0.8	6.71
2300	0.0846	0.8	7.79
2400	0.0980	1.0	8.90
2500	0.1119	1.0	10.0
2600	0.1265	1.2	11.2
2700	0.1415	1.2	12.3
2800	0.1569	1.4	13.5
2900	0.1725	1.6	14.6
3000	0.1882	1.6	15.7
3100	0.2039	1.8	16.8
3200	0.2196	2.0	17.8
3300	0.2353	2.0	18.8
3400	0.2507	2.2	19.8
3500	0.2659	2.4	20.7
3600	0.2809	2.6	21.6
3700	0.2956	2.6	22.4
3800	0.3099	2.8	23.2
3900	0.3239	3.0	24.0
4000	0.3375	3.3	24.7

TABLE II. Fitted constants for fourth-order fit to $Q(X)$, see Eq (3).

X	a_0	a_1	a_2	a_3	a_4
Total	-46.411 039	65.619 568	-33.437 298	7.485 6164	-0.599 286 97
HCN	-104.971 96	146.918 09	-75.853 937	17.337 515	-1.455 682 1
HNC	-55.717 312	79.913 945	-41.641 528	9.570 450 0	-0.797 169 03

somewhat approximately, for levels associated with J values up to 150, larger errors are likely to come from the truncation of the energy levels sums at $18\,000\text{ cm}^{-1}$ above the ground state. As the ground state of HNC is higher in energy than that of HCN, we would expect such truncations to have a larger effect on $Q(\text{HNC})$ than $Q(\text{HCN})$ for a given temperature. To estimate this truncation error we modeled the higher vibrational levels by extrapolating the energy level distribution using a quadratic formula. Using this data suggests that there is no significant contribution from these levels below 2500 K; the contributions are 0.4% at 3000 K, 1.4% at 3500 K and 3.2% at 4000 K. Above this 4000 K this error rises rapidly and is the reason why these higher temperatures are not considered here. Our partition functions and other data have not been adjusted for this systematic error.

In addition, the assignment of levels between HCN, HNC and delocalized states is another source of error. In the case of the first 302 levels it is possible that for those cases where B_v is close to 1.51 some levels have been wrongly assigned. However, it is thought unlikely that any such misassignments will affect $K(T)$ by more than 1%. A more significant source of error is likely to be the assumption that the gaps in Harris *et al.*'s assignments of states between the 302nd level and $18\,000\text{ cm}^{-1}$ have no bias to one or other form of the isotope. Although not quantifiable, it is possible that at 3000 K this could give rise errors of up to 5%.

Besides truncation errors, the other source of systematic errors in this study are associated with the use of a particular *ab initio* potential energy surface for the system. The error in many of the individual energy levels, particularly for low lying states, can be assessed by comparison with spectroscopic data. The potential surface generally performs well^{4,21} although the overestimate of ν_1 stretching mode of HNC will lead to some underestimate of $Q(\text{HNC})$. Crucial for our estimate of the equilibrium constant, K , and undoubtedly important also for the total partition function of the system, $Q(\text{Total})$, is the energy separation between the ground states of the two isomers, ΔE_{HNC} . The separation that we have used, 5185.6 cm^{-1} , is certainly not accurate to all the figures quoted. The figure is obtained from Harris *et al.*'s data which were obtained from calculations which made particular efforts to characterize the relative energy of the two isomers using the best possible *ab initio* techniques.⁴ Its value is consistent with most other estimates available at present but will undoubtedly be refined in the future.

We have estimated the effect on $K(T)$ of a 100 cm^{-1} error in the separation, ΔE_{HNC} . Assuming that the values for $Q(\text{HCN})$ and $Q(\text{HNC})$ are unaltered, changing ΔE_{HNC} by 100 cm^{-1} alters K by 5% at 2500 K which reduces to 4% at 4000 K. As T tends to zero, an error in ΔE_{HNC} gives an increasingly large relative error in the (small) $K(T)$ value.

IV. CONCLUSIONS

We have used rotation-vibration energy levels obtained from variational nuclear motion calculations using an *ab initio* surface to estimate partition functions for the HCN sys-

tem. From these data the calculation of the total partition function for the HCN system is fairly straightforward, however the calculation of separate partition functions for the HCN and HNC isomers required each level in the system to be separately labeled according to which isomer it was associated with. This task has been achieved for sufficient levels to obtain the individual partition functions and hence the temperature-dependent equilibrium constant for the HCN isomerization reaction. These calculations suggest, for example, that approaching 20% of HCN present in the atmosphere of cool carbon stars should be in the form of HNC, a molecule which, of course, has a completely different spectral signature.

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